Foundations Semester 1 Review Name:\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

1. page 450; define pure substance

2. page 451; give an example of a compound

3. page 453; give an example of a heterogeneous mixture

4. page 454; give an example of a homogeneous mixture

5. page 458; give an example of a physical property

6. page 460; give an example of a physical change

7. page 461; give an example of a chemical property

8. page 462; give an example of a chemical change

9. page 465; define the law of conservation of mass

10. page 485; define buoyancy

11. page 486; define Pascal’s principle

12. page 488; define Bernoulli’s principle

13. page 489; define viscosity

14. page 492; explain Boyle’s law

15. page 494; explain Charles’s law

16. page 513; what is the mass number and how do you calculate it?

17. page 514; define isotope

18. page 518; metals can be found where on the periodic table?

19. page 519; metalloids can be found where on the periodic table?

20. page 521; energy levels have a maximum of number of what?

21. page 570; define malleable

22. page 572; what elements are in the same family of Lithium?

23. page 582; what is unique about the noble gas family?

24. page 608; define ion

25. page 610; how does an ionic bond form?

26. page 611; how does a covalent bond form?

27. page 8; define hypothesis

28. page 9; define variable

29. page 9; define dependent variable

30. page 9; define independent variable

31. page 9; define constant

32. page 9; define control

33. page 12; define theory

34. page 12; define scientific law

35. page 15; look at table 2; write the prefixes in order from largest to smallest

36. page 22; define graph

37. page 23; write the first sentence under the heading “Line Graphs”

38. page 25; write the first sentence under the heading “Bar Graphs”

Foundation Final Review Semester 2 2013

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| --- | --- |
| 1. Popping a balloon is a \_\_\_\_\_\_\_\_\_\_\_change
 |  |
| 1. Mixtures can be separated by \_\_\_\_\_\_\_\_\_\_means
 |  |
| 1. Salsa is a \_\_\_\_\_\_\_\_\_\_\_ mixture
 |  |
| 1. Substances that cannot be broken down chemically are
 |  |
| 1. Flammable is a \_\_\_\_\_\_\_\_ property
 |  |
| 1. matter with definite volume, but no definite shape is
 |  |
| 1. matter in which particles are free to move in all directions filling a container is
 |  |
| 1. if something is heated its particles will
 |  |
| 1. An inflated balloon put in the freezer will
 |  |
| 1. a fluids resistance to flow is
 |  |
| 1. the melting point is the same as the
 |  |
| 1. The law of conservation of mass says that if you start with 10 hydrogen, you end with
 |  |
| 1. Every chemical reaction involves a change in
 |  |
| 1. The law of conservation of mass says that if you start with 10 hydrogen, you end with
 |  |
| 1. a chemical bond where atoms are transferred is a
 |  |
| 1. Why do the noble gases not form compounds
 |  |
| 1. How many hydrogen atoms are there in H4CO2OH?
 |  |
| 1. At room temperature most metals are
 |  |
| 1. An ion has 9 protons and 10 electrons its charge is
 |  |
| 1. each energy level of an atom has a maximum number of these it can hold
 |  |
| 1. Elements that are on the stair-step line in the periodic table are
 |  |
| 1. What elements are similar to Lithium because they are in the same family?
 |  |
| 1. Where do you find heavier elements in the periodic table?
 |  |
| 1. Where are metals in the periodic table?
 |  |
| 1. Where are gases in the periodic table?
 |  |

**Study these notes, expect a note check**

Introduction to science and scientific method

* yellow graphing note-taker 3-tab
* green science skills

Chapter 15 – classification of matter

– White flow chart – pure substance, mixture, etc.

* Opposites note-taker (colored) – physical property and change; chemical property and change

Chapter 16 – state of matter

 - mini book

 - temperature graph

 - pink 3 tab

 - blue 3 tab

Chapter 17,19 – Atoms

 - yellow 4 square

 - large periodic table with notes

 - small colored periodic table

 - ionic and covalent notes